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study of the relationships between amounts (number of moles or masses) of reactants and products. It is fundamental to all of chemistry, and it is essential that you understand the stoichiometry and are able to apply it to numbers of moles and masses in chemical reactions.

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molarity, divide the  
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the volume in  $\text{dm}^3$ .  
 $0.0175 \text{ mol} / 0.25 \text{ dm}^3$ .  
 $= 0.07 \text{ M}$ . So initially,  
solution EF was  
prepared so that the  
molarity of oxalic acid  
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know from question (b)

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