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Mastering Chemistry - Chapter 4. For soluble salts, the molarity of a particular ion is equal to the molarity of that compound times the subscript for that ion. For example, 1 M of AlCl₃ is 1 M in Al³⁺ and 3 M in Cl⁻. 1 M of (NH₄)₂SO₄ is 2 M in NH₄⁺ and 1 M in SO₄²⁻.

Answers To Mastering Chemistry Chapter 4

Step 4 of 6 Sodium sulfide : When dissolved in water it forms hydroxide ions therefore, it is alkaline in nature. The presence of sulfide ion can be confirmed by the following test.

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Mastering Chemistry Chapter 4 Zn (s) + MnI₂ (aq) → ZnI₂ (aq) + Mn (s) 3Hg (l) + 2Cr (NO₃)₃ (aq) → 3Hg (NO₃)₂ + 2Cr (s) 3FeBr₂ (aq) + 2Au (s) → 3Fe (s) + 2AuBr₃ (aq) 2AgNO₃ (aq) + Pb (s) → 2Ag (s) + Pb (NO₃)₂ (aq) SnCl₂ (aq) + Cu (s) → Sn (s) + CuCl₂ (aq)

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Chemistry 113 Chapter 4/5 Mastering Chemistry +- Ion Concentration? A scientist wants to make a solution of tribasic sodium phosphate, Na₃PO₄, for a laboratory experiment. How many grams of Na₃PO₄ will be needed to produce 500. mL of a solution that has a concentration of Na⁺ ions of 0.800 M ?

Chemistry help! Chemistry 113 Chapter 4/5 Mastering ...

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